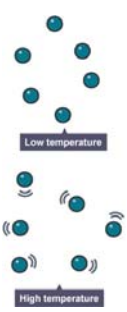


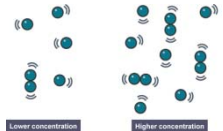
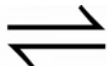
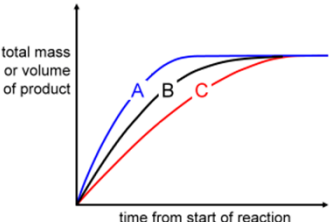

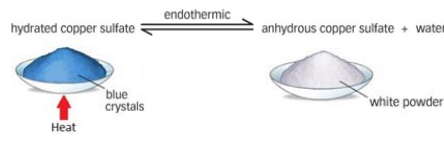

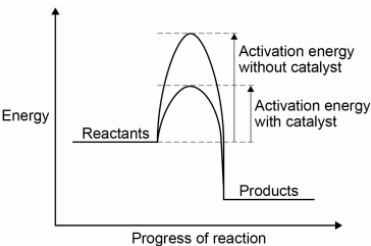


<h2 style="text-align: center;">C6 The rate and extent of chemical change</h2> <h3 style="text-align: center;">AQA Trilogy</h3>	<p>Describe how increasing the temperature can affect the rate of reaction</p> 	<p><b>RPA 11: The effect of concentration on rate of reaction</b></p> <p>Draw a labelled diagram to show how you could collect the CO<sub>2</sub> gas that is given off when you react Na<sub>2</sub>CO<sub>3</sub> with HCl</p> 	<p>Describe what happens if the temperature of a system at equilibrium is:</p> <p>Increased</p> <p>Decreased</p> 
<p>Describe how to measure a rate of reaction</p> <p>What might you expect the units of rate of reaction to be?</p>	<p>Describe how increasing the concentration or pressure can affect the rate of reaction</p> 	<p>Define the terms : activation energy</p> <p>Equilibrium</p> <p>What does this symbol mean?</p> 	<p>Describe what happens if the concentration of a system at equilibrium is:</p> <p>Increased</p> <p>Decreased</p>
<p>Which line shows the fastest rate of reaction? How would you know the reaction has finished?</p> 	<p>Describe how increasing the surface area can affect the rate of reaction</p> 	<p>If a reaction is endothermic in one direction, what can be said about the reaction in the opposite direction?</p> 	<p>Describe what happens if the pressure of a system at equilibrium is:</p> <p>Increased</p> <p>Decreased</p>
<p>By adding Mg to HCl, Amy collected 30cm<sup>3</sup> of hydrogen in the first 15 seconds. Calculate the rate of reaction and give the units</p> 	<p>Describe how using a catalyst can affect the rate of reactions</p>	<p>What does the diagram show you about the use of a catalyst?</p> 	<p>Nitrogen + Hydrogen <math>\rightleftharpoons</math> Ammonia  <math>\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3</math></p> <p>Predict what happens to the equilibrium if :</p> <p>Temperature is decreased</p> <p>Pressure is increased</p> <p>The concentration of N<sub>2</sub> is increased</p>