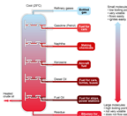

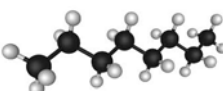

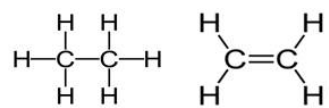


<h2 style="text-align: center;">C7 Organic Chemistry AQA Trilogy</h2>	<p>Explain how fractional distillation works to separate the different compounds in crude oil (Keywords: evaporation and condensation)</p> 	<p>Some of the hydrocarbons from crude oil are used as a feedstock for the petrochemical industry. List them below.</p> <ul style="list-style-type: none"> • • • • 	<p>How do we test for the presence of alkenes?</p> 
<p>Define the following:</p> <p>Crude oil</p> <p>Hydrocarbon</p> <p>viscous</p>	<p>As molecule length increases, what happens to the:</p> <p>Boiling point</p> <p>Viscosity</p> <p>flammability</p>	<p>Define the terms:</p> <p>Cracking</p> <p>Vaporise</p> <p>Oxidation</p>	<p>Describe how cracking is done by thermal decomposition</p>
<p>Write the general formula for an alkane</p> <p>Name the first 4 members of the alkanes</p> <ul style="list-style-type: none"> • • • • 	<p>Write a word equation for the combustion of a hydrocarbon.</p>	<p>What is the importance of cracking long chain hydrocarbons?</p> 	<p>Describe how cracking is done using a catalytic cracking</p>
<p>Draw the displayed formula for C₂H₆</p> <p>Draw the displayed formula for C₃H₈</p>	<p>Complete the balanced symbol equations</p> <p>CH₄ + O₂ →</p> <p>C₂H₆ + O₂ →</p> <p>C₃H₈ + O₂ →</p> <p>C₄H₁₀ + O₂ →</p> 	<p>How is an alkene different to an alkane?</p> 	<p>Describe how cracking is done using steam</p>